Section 4.2 The Structure of an Atom  
(pages 108–112)

This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy (page 108)

Monitoring Your Understanding  Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.

<table>
<thead>
<tr>
<th>What I Know About Atoms</th>
<th>What I Would Like to Learn</th>
<th>What I Have Learned</th>
</tr>
</thead>
<tbody>
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</table>

Properties of Subatomic Particles (pages 108–109)

1. What are three subatomic particles?
   a.  
   b.  
   c.  

2. Circle the letter that identifies a subatomic particle with a positive charge.
   a. nucleus  
   b. proton  
   c. neutron  
   d. electron  

3. Why did Chadwick conclude that the particles produced by his experiment were neutral in charge?

Comparing Subatomic Particles (pages 109–110)

4. Circle the letters of properties that vary among subatomic particles.
   a. color  
   b. mass  
   c. charge  
   d. location in the atom  

5. Circle the letter of the expression that accurately compares the masses of neutrons and protons.
   a. mass of 1 neutron = mass of 1 proton  
   b. mass of 2000 neutrons = mass of 1 proton  
   c. mass of 1 electron = mass of 1 proton  
   d. mass of 1 neutron = mass of 1 electron  

Physical Science Reading and Study Workbook  •  Chapter 4  37
Atomic Number and Mass Number (page 110)

6. Is the following sentence true or false? Two atoms of the same element can have different numbers of protons. 

7. What is an atomic number? 

8. Circle the letters that identify quantities that are always equal to an element’s atomic number.
   a. number of nuclei
   b. number of protons
   c. number of neutrons
   d. number of electrons

9. Is the following sentence true or false? Two different elements can have the same atomic number. 

10. What is the mass number of an atom? 

11. Complete the equation in the table below.

<table>
<thead>
<tr>
<th>Number of neutrons</th>
<th>Number of protons</th>
</tr>
</thead>
</table>

Isotopes (page 112)

12. Every atom of a given element has the same number of ____________ and ____________.

13. Every atom of a given element does not have the same number of ____________.

14. What are isotopes? 

15. All oxygen atoms have 8 protons. Circle the letter of the number of neutrons in an atom of oxygen-18.
   a. 8      b. 9
   c. 10     d. 18

16. Is the following sentence true or false? Isotopes of oxygen have different chemical properties. 

17. Water that contains hydrogen-2 atoms instead of hydrogen-1 atoms is called ________________.